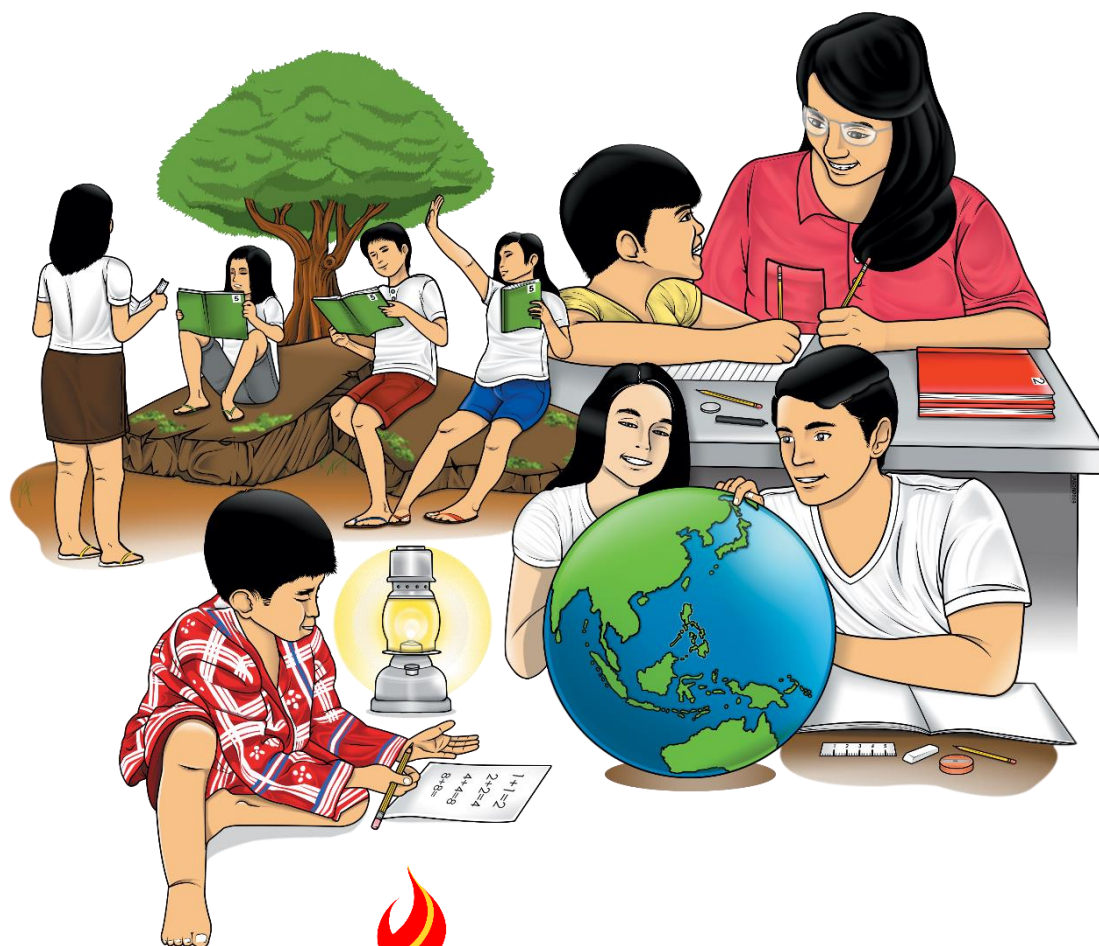


7

Science

Quarter 1 – Module 4: I Have Less, She Has Ample, He Has More, Let Us See What's In Store!



Science – Grade 7
Alternative Delivery Mode
Quarter 1 – Module 4: I Have Less, She Has Ample, He Has More,
Let Us See What’s In Store!
First Edition, 2020

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Science

**Quarter 1 – Module 4:
I Have Less,
She Has Ample,
He Has More, Let Us See
What's In Store!**

Introductory Message

For the facilitator:

Welcome to the **Science 7** Alternative Delivery Mode (ADM) Module on **Describing Saturated and Unsaturated Solutions!**

This module was collaboratively designed, developed and reviewed by educators both from public and private institutions to assist you, the teacher or facilitator in helping the learners meet the standards set by the K to 12 Curriculum while overcoming their personal, social, and economic constraints in schooling.

This learning resource hopes to engage the learners into guided and independent learning activities at their own pace and time. Furthermore, this also aims to help learners acquire the needed 21st century skills while taking into consideration their needs and circumstances.

In addition to the material in the main text, you will also see this box in the body of the module:

As a facilitator, you are expected to orient the learners on how to use this module. You also need to keep track of the learners' progress while allowing them to manage their own learning. Furthermore, you are expected to encourage and assist the learners as they do the tasks included in the module.

For the learner:

Welcome to the **Science 7** Alternative Delivery Mode (ADM) Module on **Describing Saturated and Unsaturated Solutions**.

The hand is one of the most symbolized parts of the human body. It is often used to depict skill, action and purpose. Through our hands we may learn, create and accomplish. Hence, the hand in this learning resource signifies that you, as a learner is capable and empowered to successfully achieve the relevant competencies and skills at your own pace and time. Your academic success lies in your hands!

This module was designed to provide you with fun and meaningful opportunities for guided and independent learning at your own pace and time. You will be enabled to process the contents of the learning resource while being an active learner.

This module has the following parts and corresponding icons:



What I Need to Know

This will give you an idea of the skills or competencies you are expected to learn in the module.



What I Know

This part includes an activity that aims to check what you already know about the lesson to take. If you get all the answers correct (100%), you may decide to skip this module.



What's In

This is a brief drill or review to help you link the current lesson with the previous one.



What's New

In this portion, the new lesson will be introduced to you in various ways such as a story, a song, a poem, a problem opener, an activity or a situation.



What is It

This section provides a brief discussion of the lesson. This aims to help you discover and understand new concepts and skills.



What's More

This comprises activities for independent practice to solidify your understanding and skills of the topic. You may check the answers to the exercises using the Answer Key at the end of the module.



What I Have Learned

This includes questions or blank sentence/paragraph to be filled in to process what you learned from the lesson.



What I Can Do

This section provides an activity which will help you transfer your new knowledge or skill into real life situations or concerns.



Assessment

This is a task which aims to evaluate your level of mastery in achieving the learning competency.



Additional Activities

In this portion, another activity will be given to you to enrich your knowledge or skill of the lesson learned. This also tends retention of learned concepts.



Answer Key

This contains answers to all activities in the module.

At the end of this module you will also find:

References

This is a list of all sources used in developing this module.

The following are some reminders in using this module:

1. Use the module with care. Do not put unnecessary mark/s on any part of the module. Use a separate sheet of paper in answering the exercises.
2. Don't forget to answer *What I Know* before moving on to the other activities included in the module.
3. Read the instruction carefully before doing each task.
4. Observe honesty and integrity in doing the tasks and checking your answers.
5. Finish the task at hand before proceeding to the next.
6. Return this module to your teacher/facilitator once you are through with it.

If you encounter any difficulty in answering the tasks in this module, do not hesitate to consult your teacher or facilitator. Always bear in mind that you are not alone.

We hope that through this material, you will experience meaningful learning and gain deep understanding of the relevant competencies. You can do it!

Lesson

1

SATURATED and UNSATURATED SOLUTION



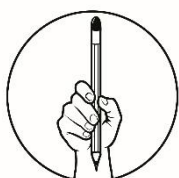
What I Need to Know

In Module 2, you found out that a solution is formed when a solute dissolves in a solvent to form a single phase that appears uniform throughout. A solution is clear and the particles are too small that they cannot be seen by the unaided eye. The particles in solution are smaller than the pores of the filter paper or the cheesecloth and so these can pass through the filter.

In Module 3, you will find out how much solute can dissolve in a given amount of solvent and find out the type of solution based on whether there is excess solute or not.

After going through this module, you are expected to:

1. Describe saturated and unsaturated solution;
2. Determine how much solid solute dissolves in each volume of water; and
3. Describe the appearance of a saturated solution.



What I Know

Directions: Read each item carefully. Write only the letter of the correct answer for each question. Use a separate sheet for your answers.

1. In a salt water solution, what substance is considered the solvent?
 - A. Salt
 - B. Water
 - C. Both are solvents
 - D. Neither substance is a solvent
2. What substance can dissolve other substance in a solution?
 - A. Solute
 - B. Mixture
 - C. Solvent
 - D. All of the above

3. What is the substance called that is being dissolved in a solution?
 - A. Solute
 - B. Mixture
 - C. Solvent
 - D. All of the Above

4. Which of the following refers to the solution that contains as much solute as can dissolve at a given temperature?
 - A. Solubility
 - B. Dilute solution
 - C. Saturated solution
 - D. Unsaturated solution

5. Which of the following refers to the solution that contains less solute than can dissolve at a given temperature?
 - A. Solubility
 - B. Dilute solution
 - C. Saturated solution
 - D. Unsaturated solution

6. You are given a 40 mL solution in a beaker. You add solute to the beaker and it dissolves completely. The solution was _____
 - A. Saturated
 - B. Unsaturated
 - C. Concentrated
 - D. Supersaturated

7. Something that can be dissolved in a solution is called _____
 - A. Colloid
 - B. Soluble
 - C. Insoluble
 - D. Suspension

8. Substance dissolved in a solution is called _____
 - A. Solute
 - B. Solvent
 - C. Solution
 - D. Concentration

9. Jessica made a pitcher of lemonade. What can she do to dilute it if she thinks it doesn't taste right?
 - A. Add water
 - B. Add sugar
 - C. Boil the lemonade
 - D. Put it in the refrigerator

10. When a solution is saturated?
 - A. Crystals form
 - B. You need to stir it more
 - C. No additional material will dissolve in it
 - D. Two materials have combined to create a clear liquid








What's In

In Grade 6, you have learned about different mixtures and their characteristics. You have done activities where you mixed a solid and a liquid or combined two different liquids. In the process of mixing, you have observed that these mixtures either form homogeneous or heterogeneous mixtures. You have seen that when all parts of the mixture have the same uniform appearance and properties, it is homogeneous.

Look at the given substances below and guess where each of the substances belongs. Answer the table below by putting a check inside the table 1 if it is homogeneous or heterogeneous mixture.

Table 1: Activity 1: Where Do I belong?

Substance	Homogeneous	Heterogeneous
1.  Gelatin		
2.  Cup of milk		
3.  Candies		
4.  Cup of coffee		
5.  Veggie Salad		

Activity 2: “Finding Solutions at Home”

List down some common solutions found at home and do the activity below

Products or Solutions Found at Home or in Stores	Characteristics
1	
2	
3	
4	
5	

Guide Questions:

1. Describe the observable characteristics of listed solutions
2. As you observe each product was described in terms of color and appearance, odor, feel, taste, and number of phases
3. Which of these products are solutions?



What's New

Activity 3: What is the Evidence that a Solution is saturated?

Perform the activity below

Materials Needed

- 6 teaspoons sugar
- 1 cup of water
- 1 measuring cup (1 cup capacity)
- 1 measuring spoon ($\frac{1}{2}$ tsp capacity)
- 2 small clear, transparent bottle
- 2 stirrers / spoon
- 1 thermometer

CAUTION: Use carefully the following laboratory instruments/kitchen materials to avoid damage and accident during the conduct of activity.

Procedure:

1. Put 20 mL (approximately 2 tablespoons) of water in a small clear transparent bottle. Add $\frac{1}{2}$ teaspoon of sugar and stir.
Q1. What is the appearance of the solutions? Write your observations.
2. To the sugar solution in step #1, add $\frac{1}{2}$ teaspoon sugar, a small portion at a time and stir the solution to dissolve the sugar. At this point, you have added 1 teaspoon sugar.

3. Add $\frac{1}{2}$ teaspoon of sugar to the sugar solution in step #2 and stir the solution. At this point, you have added one and $\frac{1}{2}$ teaspoons of sugar.
4. Continue adding $\frac{1}{2}$ teaspoon sugar to the same cup until the added sugar no longer dissolves.

Q2. How many teaspoons of sugar have you added until the sugar no longer dissolves? _____ Teaspoons

Note: In this step, you will observe that there is already excess sugar which did not dissolve.

Q3. So, how many teaspoons of sugar dissolved completely in 20 mL of water? _____ Teaspoons

Note: This is now the maximum amount of sugar that will completely dissolve in 20 mL of water.



What is It

The substances that make up a homogeneous solution are called components of the solution. These components are called solvent and a solute. What is solvent? What is solute?

Solvent it is a component of a solution which dissolves the other component in itself. It institutes the larger component of the solution. For example, water is a solvent that dissolves solid substance like sugar.

Solute it is the component of the solution which dissolves in the solvent. It has the lesser component of the solution. For example, sugar is a solute that dissolves in water.

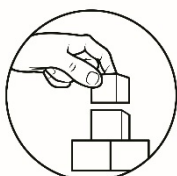
In Activity 3, you observed that the appearance of solution containing less amount of solute (sugar) was clear or transparent. When sugar is dissolved in water, the particles of sugar gets between the spaces of the particles of water and creates a single phase of solution. However, when you slowly add more amount of sugar into a solution and stir it, you will observe that the solution reaches already the point at which it cannot dissolve more solute and the sugar you add sinks to the bottom in solid form.

You have observed that there is a maximum amount of solute like sugar that can dissolve in a given amount of solvent like water at a certain temperature. The process took place between sugar and water is called the **solubility** of the solute.

The solution that contains the maximum amount of solute dissolved by a given amount of solvent is called a **saturated solution**. The presence of an excess solid which can no longer dissolve is evidence that the solution is saturated. A solution is **unsaturated** when it contains less solute than the maximum amount it can dissolve at a given temperature.

Guide Question:

1. What is the difference between the Unsaturated and saturated solutions?
2. How do you know a solution is saturated?
3. What happens when a solution becomes saturated?
4. How many grams or teaspoons of sugar dissolved in 20 ml of water to form saturated solution?



What's More

Study the table below and answer the following questions.

Note: The 35 grams of table salt will form saturated solution in 100 ml of water.

Amount of Table salt in grams (g)	Amount of water in milliliter (ml)
1. 20	100
2. 35	100
3. 70	100

Guide Questions:

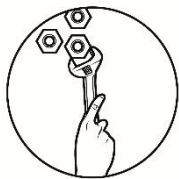
1. In which amount of table salt and water will form an unsaturated solution?
2. How many grams of table salt will dissolve to water to form a saturated solution?
3. Which of the following will form a supersaturated solution?



What I Have Learned

Complete the statements below to express what you had learned from the lesson.

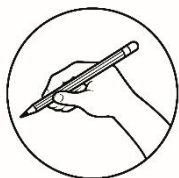
1. A maximum amount of solute that can dissolve in a given amount of solvent at a certain temperature is called the _____ of the solute.
2. The solution that contains the maximum amount of solute dissolved by a given amount of solvent is called a _____.
3. The presence of an excess solid which can no longer dissolve is evidence that the solution is _____.
4. A solution _____ when it contains less solute than the maximum amount it can dissolve at a given temperature.



What I Can Do

List down some examples of solutions that we need to prepare/make in the form of Unsaturated and Saturated solution.

Unsaturated	Saturated



Assessment

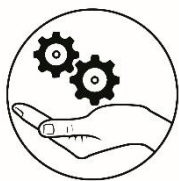
Directions: Read each item carefully. Write only the letter of the correct answer for each question. Use a separate sheet for your answers.

1. You are given a 40 mL solution in a beaker. You add solute to the beaker and you observed some particles did not dissolve. What solutions is it?
 - A. Saturated
 - B. Unsaturated
 - C. Concentrated
 - D. Supersaturated
2. What do you call a substance that dissolved in another substance which is in greater amount?
 - A. Solute
 - B. Solvent
 - C. Solute and Solvent
 - D. Neither Solute nor Solvent
3. What do you call a substance dissolved in any solution?
 - A. Solute
 - B. Solvent
 - C. Solute and Solvent
 - D. Neither Solute nor Solvent

4. What you can do if you add more amount of sugar in a cup of your hot milk and it taste very sweet?
 - A. Add water
 - B. Add sugar
 - C. Mix the milk solution well
 - D. Put it in the refrigerator for an hour
5. When a solution is saturated?
 - A. Crystals form
 - B. You need to stir it more
 - C. No additional material will dissolve in it
 - D. Two materials have combined to create a clear liquid
6. How will you prepare an unsaturated solution?
 - A. Freeze the mixture
 - B. Stir the powder in the liquid
 - C. Add less amount of powder to the liquid
 - D. Add more amount of solute in a lower amount of solvent
7. To make a solute dissolve more quickly in a solvent, which would you do?
 - A. Stir it
 - B. Do not stir the solution
 - C. Let the solute settle down
 - D. Nothing to do with the solute
8. How will you define solubility?
 - A. Lack of polarity of molecules
 - B. Amount of polarity of molecules
 - C. Ability of a solvent to dissolve in a solute
 - D. Ability of a solute to dissolve in a solvent
9. Which refers to greater amount needed in dissolving mixtures?
 - A. Solute
 - B. Solvent
 - C. Solution
 - D. Hydration
10. Which of these factors will cause more sugar to dissolve in a saturated sugar solution?

- I. Add more sugar while stirring
- II. Add more sugar and heat the solution
- III. Add more sugar to the heated solution
- IV. Add more sugar and cool down the solution

- A. I, II, III only
- B. I and II only
- C. II and III only
- D. I, III, IV only



Additional Activities

Saturated solutions aren't just for science. Saturated solutions and solubility play an important role in our lives, especially in the kitchen. In this home kitchen experiment, we will be examining how temperature relates to solubility.

CAUTION: Please do the activity seriously to avoid some problems.
Be careful in handling hot water!

Directions:

1. Prepare two large cups and place it on the plane table.
2. Add hot water in one of the cup and cold water in the other cup.
3. Next, spoon by spoon add as much sugar as you can if it keeps dissolving. Do this to the hot water first to avoid it cooling, and then repeat with the cold water.
4. Record how many spoonful of sugar you can add in each cup until it is completely dissolved.

Temperature	Spoonful of Sugar
Cold	
Hot	

Question:

Which temperature had a greater solubility and how did you know?
Include observations from your experiment.

Lesson

2

SUPERSATURATED SOLUTION



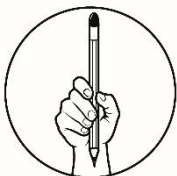
What I Need to Know

In lesson 1, you found out that you need to dissolve a given amount of a substance (solute) in the required amount of water (Solvent) to produce a saturated and unsaturated solution.

Here in Lesson 2, you will understand supersaturated solution

After going through this lesson, you are expected to:

1. Define supersaturated solution
2. Compare and contrast saturated and supersaturated
3. Cite the importance of supersaturated solution



What I Know

Directions: Read each item carefully. Write only the letter of the correct answer for each question. Use a separate sheet for your answers.

1. Which of the following substances is most soluble in water?
 - A. Flour
 - B. Table sugar
 - C. Cooking oil
 - D. Baking soda

2. Which statement is true of any saturated solution at a given temperature?
 - A. No more solute will dissolve in the solution.
 - B. Adding more solute will increase the saturation
 - C. You can dissolve more solute if you stir the solution.
 - D. One liter of the solution contains 2000 grams of solute.

3. Which of the following refers to the maximum amount of solute that can dissolve in a given amount of solvent at a certain temperature?
 - A. solubility
 - B. dilute solution
 - C. saturated solution
 - D. Unsaturated solution

4. Which of these solutions has more solute than it can hold?
 - A. Saturated
 - B. Suspension
 - C. Unsaturated
 - D. Supersaturated

5. Supersaturated solution is one with_____?
 - A. Greater amount of solvent
 - B. Less solute than the solvent
 - C. Less solvent than the solute
 - D. Equal amounts of solute and solvent

6. What is something that can be dissolved in a solution called?
 - A. Colloid
 - B. Soluble
 - C. Insoluble
 - D. Suspension

7. What do you call a substance dissolved in a solution and contains greater amount of solute than the other component?
 - A. Saturated
 - B. Suspension
 - C. Unsaturated
 - D. Supersaturated

8. Jessica wanted to ferment a fish. What type of salt solution she needs to prepare in order to conduct fermentation?
 - A. Saturated
 - B. Suspension
 - C. Unsaturated
 - D. Supersaturated

9. When does a solution become a saturated?
 - A. Crystals form
 - B. You need to stir it more
 - C. No additional material will dissolve in it
 - D. Two materials have combined to create a clear liquid

10. A powder is about to be poured into the liquid. Which of the following should be done to make this powder dissolve faster?
- Freeze the mixture
 - Stirs the powder in the liquid
 - Add more powder to the liquid
 - Store the mixture in a dark place



What's In

What have you learned from the previous lesson? Let's try to check your prior-knowledge.

Procedures:

- Arrange the letters of the following scrambled words in column A to make it correct.
- Define each word by matching it to the given sentences in column B.
- Write your answer in column C.

Term (A)	Meaning(B)	Answer (C)
LUBILSOITY	1. a solution in which the amount of solute is equal to the solute's solubility at a given volume and temperature	
TEDSATURAUN TIONSOLU	2. is the ability of a substance to be dissolved in another substance at a specific temperature and pressure.	
RATEDSATU LUTISOON	3. a solution in which the amount of solute is less than the solute's solubility at a given volume and temperature	

Look! What is in the picture below?



What will happen to the solutions if you will add more amounts of solutes like sugar in a cup of hot coffee or cocoa powder in your hot cocoa drink? _____



What's New

Activity 3: Making Supersaturated Solution

Materials Needed

- Sugar
- Water
- Beaker (250 mL) /Plastic cups
- Stirring Rod/Spoon
- Alcohol lamp/Candle
- Tripod/Any supporting cooking material
- Wire gauze/Screen (15 cm²)
- Match stick

CAUTION: Use carefully the following laboratory instruments/kitchen materials to avoid damage and accident during the conduct of activity. Be careful in doing the activity especially if fire is involved. Make sure the presence of any adult companions while doing this activity.

Procedures:

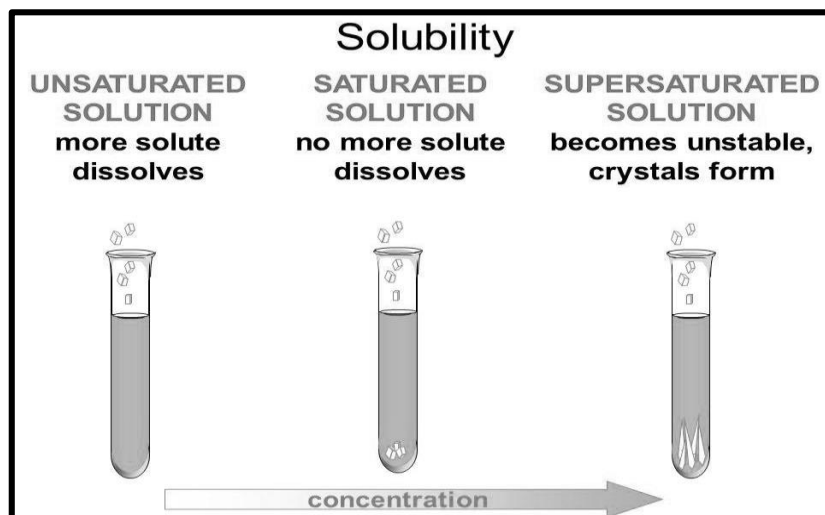
1. GET YOUR SAFETY GOOGLES AND APRONS ON!
2. Prepare a 250 ml beaker and stirring rod.
3. Fill the beaker with 100 ml of tap water.
4. Slowly add 100 grams of sugar to 100 mL water and stir it.
Q1. Observe and describe what happens to the solution.
5. Next, add another 100 grams of sugar to the sugar solution in step # 4.
Q2. Do the particles of sugar still dissolved in water?
6. Heat your solution of at least 80 °C and continue adding 100 grams to the same sugar into your solution.
Q3. This time, what had you observed from the solution?
Q4. Be specific about the particles, did they still dissolve in water? Did the water change in color?
7. Stop boiling your solution. Add another 60 grams of sugar. Stir and observe what will happen?
Q5. Describe your solution.
Q6. What do you call the solution that contains more dissolved solute than a saturated solution does?
8. Let your substance cool down its temperature up to 25 °C
Q7. What have you observed?
Q8. Are crystals formed in your solution after cooling it?



What is It

When you add more amount of solute in a solution and then you will observe some particles are left at the bottom of the solution, at this point the solution is saturated. It is a solution in which the amount of solute is equal to the solute's solubility at a given volume and temperature.

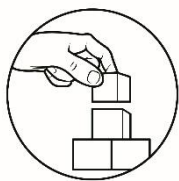
During heating of solution, you continue added solute in a solution and it still dissolved. At this portion the solution is saturated at a higher temperature. Allowing the solution to cool down at normal temperature and no solute crystallizes, the solution is **Supersaturated Solution**. It is a solution in which the amount of solute is greater than the solute's solubility at a given volume and temperature. This state is unstable, which by slight agitation causes precipitation. In this process, some of the solute will come out of the solution. Once precipitation occurs, the end result is a saturated solution.



<https://www.google.com.ph/search?q=comparison+between+saturated+supersaturated&sxsrf=ALeKk00JReTyZducFmT9pcwwsNNgkQCnmw>

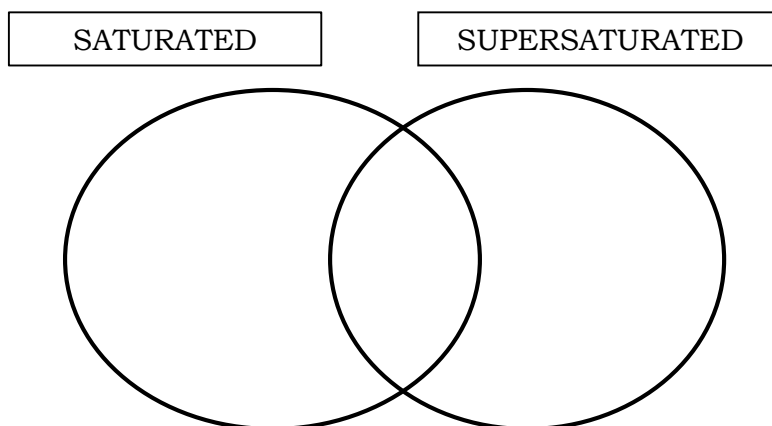
Question:

What do you think are the differences and similarities between saturated, unsaturated and supersaturated solution?



What's More

Using the diagram below, compare and contrast between Saturated and Unsaturated solutions. Write the similarities at the center of the diagram and their differences in both sides of diagram.



What I Have Learned

Part A: Classify whether the solution described is saturated or unsaturated or supersaturated.

Solubility (grams solute per 100 mL solvent) at 25°C	Solution at _____	Type of Solution
36	36 grams in 100 mL solvent at 25°C	
80	80 grams in 100 mL solvent at 25°C	
40	80 grams in 100 mL solvent at 50°C	
50	100 grams in 200 mL solvent at 25°C	
100	150 grams in 100 mL solvent at 50°C	

Part B: Read the statement inside the box and identify if the following solutions are saturated, unsaturated, or supersaturated solutions below.

A solution containing 200 grams of sucrose (or table sugar) in 100 mL of water at 25 °C is a saturated solution.

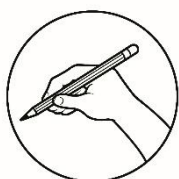
1. A solution made by dissolving 200g of the same sugar in 100mL of water at 25°C.
2. An additional 100 g of the same sugar was added to the previous solution in 100 mL of water at 25 °C.
3. The saturated solution is heated to 100 °C. All the solutes are dissolved in water.
4. Cooling down the solution to 25 °C and no crystals were formed.



What I Can Do

Congratulations for the job well done! Now to develop your writing skill, kindly make/compose a poem related to Supersaturated Solution. Write your answer in a separate sheet of paper.

TITLE



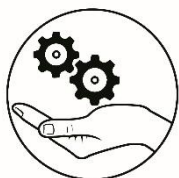
Assessment

Directions: Read each item carefully. Write only the letter of the correct answer for each question. Use a separate sheet for your answers.

1. How will you identify a saturated solution?
 - A. Less solvent than solute
 - B. Less solute than solvent
 - C. Equal amounts of solute and solvent
 - D. Less amount of both solute and solvent
2. Which statements are true of any saturated solution at a given temperature?
 - A. No more solute will dissolve in the solution.
 - B. Adding more solute will increase the saturation
 - C. You can dissolve more solute if you stir the solution.
 - D. One liter of the solution contains 2000 grams of solute.
3. What is solubility refers to its ability?
 - A. Solute to dissolve a solvent
 - B. Solute to dissolve in a solvent
 - C. Solvent to dissolve in a solute
 - D. Solvent and solute to dissolve each other

4. Which of the following refers to the solution that contains the maximum amount of solute dissolved by a given amount of solvent?
 - A. Solubility
 - B. Saturated solution
 - C. Unsaturated solution
 - D. Supersaturated solution
5. Which of the following refers to the solution that contains less amount of solute than can dissolve at a given temperature?
 - A. Solubility
 - B. Saturated solution
 - C. Unsaturated solution
 - D. Supersaturated solution
6. You are given a 30 mL solution in a beaker. You add solute to the beaker and it dissolves completely. What is the solution?
 - A. Solubility
 - B. Saturated solution
 - C. Unsaturated solution
 - D. Supersaturated solution
7. Which of the following are made up of solutes and solvents?
 - A. Colloid
 - B. Mixtures
 - C. Solutions
 - D. Suspension
8. What is the solution that contains more solute than a saturated solution under the same conditions?
 - A. Solution
 - B. Saturated
 - C. Unsaturated
 - D. Supersaturated
9. Why increasing temperature of the solvent will speed up the dissolving process? because it brings _____.
 - A. more solute to crystallize
 - B. less solute molecules to the solvent
 - C. fresh solvent into contact with more solute
 - D. more solvent molecules to collide with the solute
10. When does a solution become saturated?
 - A. Crystals form
 - B. You need to stir it more
 - C. No additional material will dissolve in it
 - D. Two materials have combined to create a clear liquid
11. Which of the following describes a solvent?
 - A. It's a metal molecule
 - B. Another word for solution
 - C. A thing that makes drinks turn colors
 - D. The liquid in which a solute is dissolved to form a solution

12. To make a solute dissolve more quickly in a solvent which would you do?
- Stir it in cold water
 - Stir it in warm water
 - Let the solute settle down
 - Nothing to do with the solute
13. Which of these does the dissolving that is usually presented in greater amounts?
- Solute
 - Solvent
 - Solution
 - Hydration
14. How does the solubility of a solid change when the temperature of the liquid solvent is increased?
- The solubility increases
 - The solubility decreases
 - There is no change in the solubility
 - The change in the solubility is unpredictable
15. A saturated solution is made by dissolving 36.8g of a solid in 200mL of water. A second solution is made by dissolving 19.1 g of the same solid in 100mL of water. How the solution would be classified?
- Unsaturated
 - Saturated
 - Supersaturated
 - Hyper saturated



Additional Activities

Prepare the following materials:

- Balance/Digital weighing scale
- Beaker/Plastic cup
- Thermometer
- Stirring rod/Spoon
- Graduated Cylinder/ Any measuring cup
- Funnel
- Sodium Chloride (Table Salt)

Procedures:

1. Make your own procedure for determining the solubility of sodium chloride in water at 30°C.
2. Create a supersaturated solution using table salt as your solute
3. Make an observation regarding the result of your activity.



Answer Key

LESSON 1

What I Know
1.A
2.C
3.A
4.C
5.D
6.
7.B
8.A
9.B
10.C

Assessment
1.A
2.B
3.A
4.A
5.C
6.C
7.A
8.D
9.B
10.C

LESSON 2

What I Know
1.B
2.A
3.C
4.D
5.C
6.B
7.D
8.D
9.C
10.B

Assessment
1.C
2.A
3.B
4.B
5.C
6.C
7.C
8.D
9.D
10.A
11.D
12.B
13.B
14.A
15.C

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